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18 States of Matter **Three**
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13) Flashcards | Quizlet*

You are already familiar with the three common states of matter: solid, liquid, and gas. Solid objects litter the room around you.

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Guide Answers
For example, you can easily recognize the shape of your desk; you know that your backpack cannot hold seven textbooks. You encounter liquids throughout the day as yo u

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Chapter 13: States of Matter

Chapter 13 States of Matter
137 SECTION 13.1 THE
NATURE OF GASES (pages
385–389) This section
introduces the kinetic
theory and describes how it
applies to gases. It defines

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Guide Answers and explains how temperature is related to the kinetic energy of the particles of a substance. Kinetic Theory and a Model for Gases (pages 385–386) 1.

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There are three states of matter that we will learn about in this chapter. (If you want to learn about more states of matter, I can refer you to somebody.) Those three states are

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Guide Answers
Solid, liquid, and gas. ?

These three states are quite different. The main difference is in their particles.

*Chapter 13: States of Matter
- Chemistry by Anna*

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Chapter 13 States of
Matter137 SECTION 13.1 THE
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introduces the kinetic
theory and describes how it
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NATURE OF GASES (pages 385–389) This section introduces the kinetic theory and describes how it applies to gases. It defines gas pressure and explains how temperature is related to the kinetic energy of the

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particles of a substance.

*Chapter 13 States Of Matter
Worksheet*

Chapter 13: States of
Matter. STUDY. PLAY. Kinetic
Molecular Theory. Explains
the properties of gases in

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Guide Answers
terms of the energy, size,
and motion of their
particles. Elastic
Collision. Describes a
collision in which kinetic
energy may be transferred
between colliding particles
but the total kinetic energy

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of the two particles remains
the same.

*Chapter 13: States of Matter
Flashcards / Quizlet*

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answers to Chapter 13 -

States of Matter - 13.1 The

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Guide of Answers - 13.1

Lesson Check - Page 424 8

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Chapter 13 - States of

Matter ...

all matter consists of tiny
particles that are
constantly in motion What

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are the three assumptions of the kinetic theory as it applies to gases? -The particles in a gas are considered to be small, hard spheres with an insignificant volume. -The motion of the particles in a

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gas are rapid, constant, and random.

*Chapter 13: States of Matter
Flashcards | Quizlet*

The Sustainable Development Goals are a call for action by all countries - poor,

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rich and middle-income - to
Guide Answers
promote prosperity while
protecting the planet. They
recognize that ending
poverty ...

*United Nations Sustainable
Development - 17 Goals to*

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Chemistry & You; 13.2 The
Nature of Liquids - 13.2
Lesson Check; 13.3 The
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13.4 Changes of State -

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Matter ...*

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13.1 The Nature of Gases -
Sample Problem 13.1; 13.1
The Nature of Gases - 13.1
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Chemistry & You; 13.2 The
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Lesson Check; 13.3 The

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Guide of Solids - Chemistry
& You; 13.3 The Nature of
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Guide Answers Chapter 13 States
of Matter 2 Kinetic Theory
as Applied to Gases
Fundamental assumptions
about gases. The particles
in a gas are considered to
be small, hard spheres with
an insignificant volume.

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Guide Answers
Between particles in a gas there is empty space. No attractive or repulsive forces exist between the particles. 3

*Chapter 13 States Of Matter
Worksheet*

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Guide Answers
The attacks the USPS
continues to face are not
just attacks on the postal
service but attacks on Black
lives. To defund the USPS
would be to deny future
generations this opportunity
and dishonor the legacy of

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Black postal workers. Now,
we're taking this matter
into our own hands by
writing and sending
#BlackLoveLetters through
USPS...

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